

AMENDMENTS TO THE SPECIFICATION:

Please replace the paragraph at page 1, after the title and before line 1, with the following amended paragraph:

This application is a divisional of U.S. Patent Application Serial No. 10/191,479 filed July 10, 2002, now Patent No. 6,657,072, which in turn is a divisional of U.S. Patent Application Serial No. 09/613,293 filed July 10, 2000, now Patent No. 6,441,216.

Please replace the paragraph beginning at line 14 of page 1, with the following amended paragraph:

These batteries consist of cathode, anode, separator and a nonaqueous electrolyte. The cathodes used are typically $\text{Li}(\text{MnMe}_z)_2\text{O}_4$, $\text{Li}(\text{CoMe}_z)\text{O}_2$, $\text{Li}(\text{CoNi}_x\text{Me}_z)\text{O}_2$ or other lithium intercalation and insertion compounds. Anodes can consist of lithium metal, carbon materials, graphite, graphitic carbon materials or other lithium intercalation and insertion compounds or alloy compounds. The electrolyte used is in the form of solutions comprising lithium salts such as LiPF_6 , LiBF_4 , LiClO_4 , LiAsF_6 , LiCF_3SO_3 , $\text{LiN}(\text{CF}_3\text{SO}_2)_2$ or $\text{LiC}(\text{CF}_3\text{SO}_2)_3$ and

LiPF_6 , $(\text{Li}^+[\text{PF}_6^-]$, lithium hexafluorophosphate), LiBF_4 ($\text{Li}^+[\text{BF}_4^-]$, lithium tetrafluoroborate), LiClO_4 ($\text{Li}^+[\text{ClO}_4^-]$, lithium perchlorate), LiAsF_6 ($\text{Li}^+[\text{AsF}_6^-]$, lithium hexafluoroarsenate), $\text{Li}[\text{SO}_3\text{CF}_3]$ ($\text{Li}^+[\text{SO}_3\text{CF}_3^-]$ lithium trifluoromethanesulfonate), $\text{Li}[\text{N}(\text{SO}_2\text{CF}_3)_2]$ ($\text{Li}^+[\text{N}(\text{SO}_2\text{CF}_3)_2^-]$, lithium bis(trifluoromethanesulfonyl)imide) $\text{Li}[\text{C}(\text{SO}_2\text{CF}_3)_3]$ ($\text{Li}^+[\text{C}(\text{SO}_2\text{CF}_3)_3^-]$, lithium tris(trifluoromethanesulfonyl)methide), and mixtures of these in aprotic solvents.

Please replace the paragraph beginning at line 25 of page 1 with the following amended paragraph:

The standard conducting lithium salts have various drawbacks. Some conducting salts having low cycling yields (e.g. LiBF_4 $\text{Li}^+[\text{BF}_4^-]$). Other conducting salts have low thermal stability (e.g. LiPF_6 $\text{Li}^+[\text{PF}_6^-]$), and yet other conducting salts are not particularly suitable because of their toxicity and poor environmental safety (e.g. LiAsF_6 $\text{Li}^+[\text{AsF}_6^-]$).

Please replace the paragraph beginning at line 25 of page 6, with the following amended paragraph:

Alternatively, the lithium borate complexes can also be used in electrolytes comprising conventional conducting salts. Suitable, for example, are electrolytes comprising conducting salts selected from the group LiPF₆, LiBF₄, LiClO₄, LiAsF₆, LiCF₃SO₃, LiN(CF₃SO₂)₂ or LiC(CF₃SO₂)₃, LiPF₆ (Li⁺ [PF₆]⁻, lithium hexafluorophosphate), LiBF₄ (Li⁺ [BF₄]⁻, lithium tetrafluoroborate), LiClO₄ (Li⁺ [ClO₄]⁻, lithium perchlorate), LiAsF₆ (Li⁺ [AsF₆]⁻, lithium hexafluoroarsenate), Li[SO₃CF₃]₂ (Li⁺ [SO₃CF₃]⁻, lithium bis(trifluoromethanesulfonyl)imide), Li[C(SO₂CF₃)₃] (Li⁺ [C(SO₂CF₃)₃]⁻, lithium tris(trifluoromethanesulfonyl)methide), and mixtures of these. The electrolytes may also comprise organic isocyanates (DE 199 44 603) to reduce the water content. Equally, the electrolytes can comprise organic alkali metal salts (DE 199 10 968) as an additive. A suitable example is that of alkali metal borates of the general formula.